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Well it gives you the $[\text{del}]\text{Hmol}^{-1}$ for the reaction as -118 KJmol^{-1} , and from the equation you can see that 2 moles of HCl react with $\text{Ba}(\text{OH})_2$, all you gave to do to find the energy change is to find out how many moles of each you have in the question.
moles of HCL = $c * v = 0.5 * 0.1 = 0.05$ moles
moles of $\text{Ba}(\text{OH})_2 = c * v = 0.5 * 0.3 = 0.15$ moles
Since the ratio is 2:1, you can see that there ...

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